



Dimerization of functional pyrroloindolizines for the synthesis of complex myrmecarin alkaloids

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ABSTRACT

The union of functionalized pyrroloindolizines for the synthesis of heterodimeric products relevant to myrmecarin alkaloids is described. Design and synthesis of tricyclic substrates and new methods for their union enable the investigation of late-stage cyclopentannulation strategies. The rapid assembly of dimeric structures using unique modes of pyrroloindolizine reactivity presents a concise approach to the dimeric myrmecarins and relevant derivatives.

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1. Introduction

The myrmecarins are a family of exceedingly air-sensitive alkaloids isolated from the poison gland of the African ant species *Myrmecaria opaciventris* (Fig. 1).¹ While these alkaloids are responsible for the paralytic activity of the secretion, their mechanism of action remains unknown.² The relative stereochemistries of myrmecarin 430A (**4**)^{1b} and myrmecarin 663 (**6**)^{1c} have been elucidated through a series of spectroscopic studies by Schröder and co-workers, whereas the extreme fragility and limited quantities of myrmecarin 645 (**5**)^{1c} have precluded stereochemical assignment. The fascinating structures and marked sensitivity of these alkaloids prompted us to develop methods for concise assembly of the complex members. We envisioned that a biogenetically inspired dimerization of activated pyrroloindolizine derivatives may provide rapid access to the highly sensitive complex structures.³

Pyrroloindolizine structures with a Lewis base substituent at C8 are predisposed to generate azafulvenium ions upon electrophilic activation (Scheme 1). This reactivity enables a range of activated derivatives to undergo addition by neutral vinyl pyrroloindolizines to afford corresponding dimeric or heterodimeric compounds.⁴ Use of C8-heteroatom substituted vinyl pyrroloindolizine nucleophiles

provides hexacyclic derivatives bearing a C1 functional group relevant to late-stage cyclopentannulation chemistries. We envisioned that strategic design of the dimerization partners would enable us to access air and acid stable hexacyclic structures with enhanced capacity for functional group manipulation at C1. With

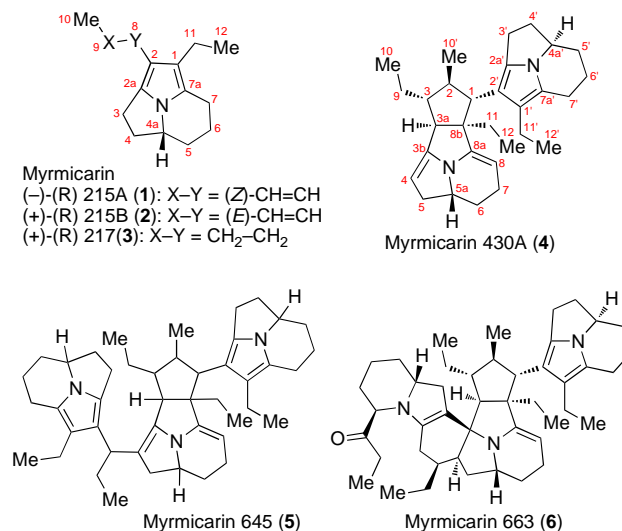
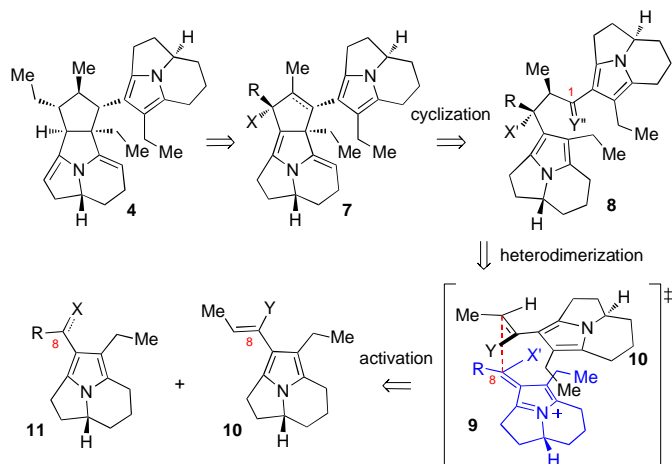


Figure 1. Representative myrmecarin alkaloids.

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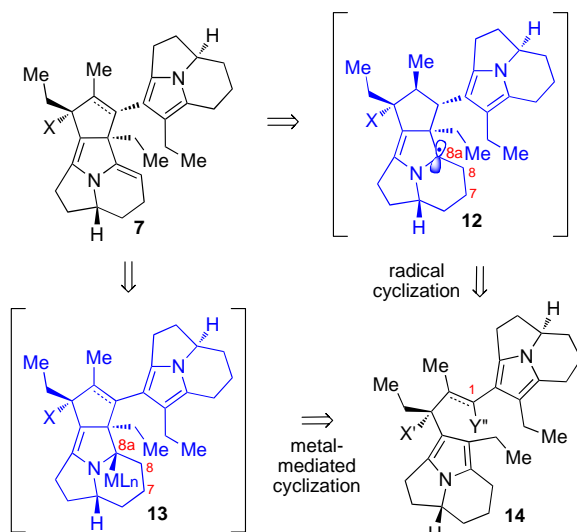
this in mind, we embarked on the development of versatile methods for fragment assembly to allow rapid investigation of alternative radical,⁵ metal-mediated,⁶ and electrophilic cyclization methods. Herein we describe the synthesis of functionalized pyrroloindolizine dimers and our studies on the unique reactivity of these structures for the synthesis of complex myrmicarins.



Scheme 1. Retrosynthetic analysis of myrmicarin 430A (4).

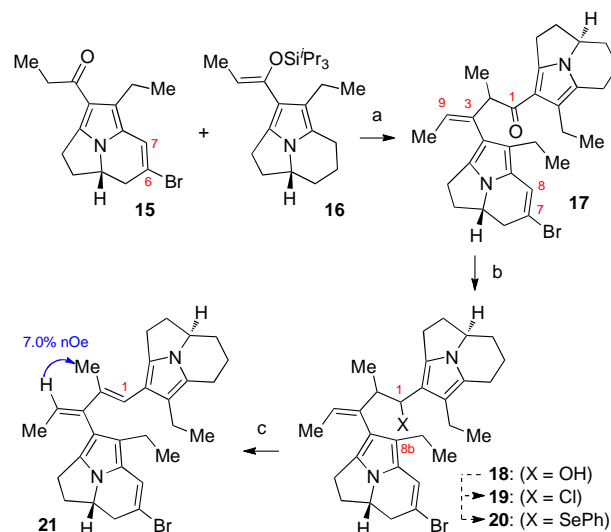
2. Heterodimerization of strategically functionalized pyrroloindolizines

To facilitate various prospective modes of cyclization we sought tricyclic subunits that would stabilize a radical or transition metal substituent at C8a (Scheme 2). We reasoned that conjugation to an alkene bearing an appropriate substituent at C7 would confer enhanced stability. Use of a corresponding functionalized tricyclic ketone as an electrophile (**11**, Scheme 1) would introduce this substituted alkene in a dimeric structure. With these considerations, we identified readily available tricycle **15**⁷ (Scheme 3) as an appropriate heterodimerization substrate to introduce the desired alkene and vinyl bromide substituent as a versatile handle for modification.



Scheme 2. Possible modes of cyclization of heterodimer **12**.

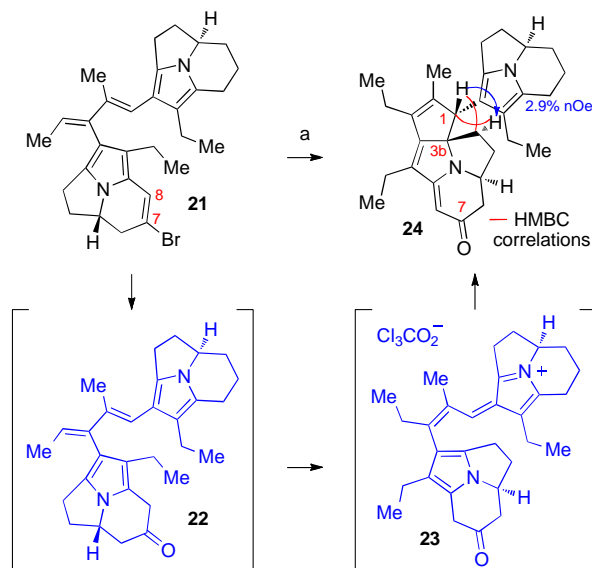
To effect heterodimerization employing **15** as an electrophile we focused on application of our previously reported trifluoromethanesulfonic anhydride-mediated activation of **15** and π -nucleophilic addition of silyl enol ether **16**.^{3a} At the outset, we



Scheme 3. Synthesis of hexacycle **18** and dehydration product **21**. Conditions: (a) TF_2O , 2,6-di-*tert*-butyl-4-methylpyridine, CH_2Cl_2 , -78°C , 2 h, 78%. (b) DIBAL-H, CH_2Cl_2 , -78°C , 2 h, 100%. (c) AcOH , C_6D_6 , 23°C , 1.8 h, 68%.

noted that appendage of the electron-withdrawing alkene to the pyrrole π system may reduce the capacity of **15** to undergo electrophilic activation, potentially diminishing the efficiency of the reaction. Gratifyingly, portion-wise addition of trifluoromethanesulfonic anhydride to a dichloromethane solution of ketone **15** and silyl enol ether **16** (1:1) in the presence of 2,6-di-*tert*-butyl-4-methylpyridine smoothly provided heterodimer **17**. As noted with previous substrates, completely regioselective elimination of trifluoromethanesulfonic acid proceeded via deprotonation at the less hindered site to provide the C3–C9 alkene. Neither the yield nor the rate of this reaction was significantly influenced by the presence of the bromoalkene in **15**, underscoring the versatility of this protocol with regard to the structure of the ketone electrophile.^{3a} To investigate formation of the C1–C8b bond we targeted derivatives that would provide a pseudobenzyl C1 radical upon activation.⁵ Examination of number of hydride reducing agents revealed that efficient reduction of the vinyllogous amide carbonyl at C1 in the presence of the C7 bromoalkene in **17** was achieved using diisobutylaluminum hydride in dichloromethane at -78°C (Scheme 3). As noted in structures lacking the C7–C8 bromoalkene, **18** exhibited marked sensitivity to conditions involving activation of the C1 alcohol. Treatment of **18** with Lewis or Brønsted acids in order to promote substitution of the C1 alcohol resulted in facile elimination to yield the exceedingly air-sensitive diene **21**. Protocols to introduce C1 halide or selenide derivatives aimed at avoiding electrophilic ionization of the C1 alcohol resulted in decomposition of **18**.

Interestingly, no cyclopentannulation byproducts arising from intramolecular alkylation at C3b were observed in these studies. By contrast, corresponding structures lacking a C7–C8 bromoalkene undergo rapid alkylation at C3b without visible accumulation of hexacyclic alkene intermediates.⁸ Pursuing this observation, we found that treatment of a benzene solution of **21** containing trace amounts of water with one equivalent of trichloroacetic acid effected conversion to heptacycle **24** (Scheme 4). Full structural characterization of this air-sensitive compound using a combination of two-dimensional NMR techniques revealed that the vinyl bromide **21** had undergone conversion to the corresponding C7 ketone **24**. The absence of heptacyclic products possessing the C7–C8 bromoalkene and failure of **18** to cyclize upon activation of the C1 alcohol with acetic acid suggest that a stronger acid was needed for the hydrolysis of the bromoalkene to enable C3b alkylation.

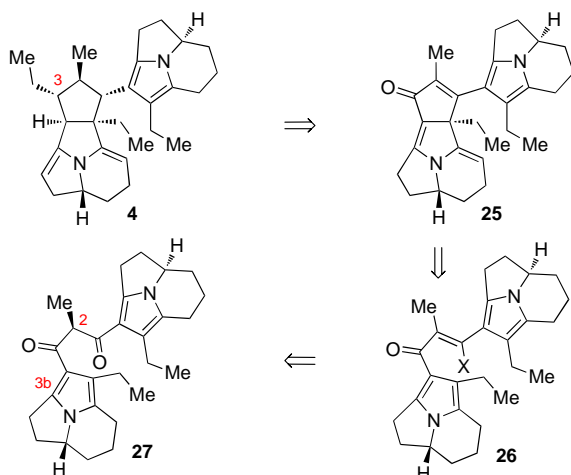


Scheme 4. Acid-promoted hydrolysis and cyclization of **21**. Conditions: (a) $\text{Cl}_3\text{CCO}_2\text{H}$, C_6H_6 , 23 °C, 1.8 h, 48%.

Importantly, the suppression of this alkylation by the presence of the C7–C8 bromoalkene demonstrated that strategic structural modification of this pyrroloindolizine subunit could be used to prevent undesired alkylation. While these studies demonstrated the applicability of our trifluoromethanesulfonic anhydride activation protocol to structurally modified pyrroloindolizine electrophiles, manipulation of the resulting heterodimers remained challenging. The sensitivity of hexacyclic compounds bearing ionizable C1 substituents motivated us to consider hexacyclic structures that would show enhanced stability and propensity for functionalization to alternative cyclization substrates.

3. Preparation and functionalization of stable heterodimers

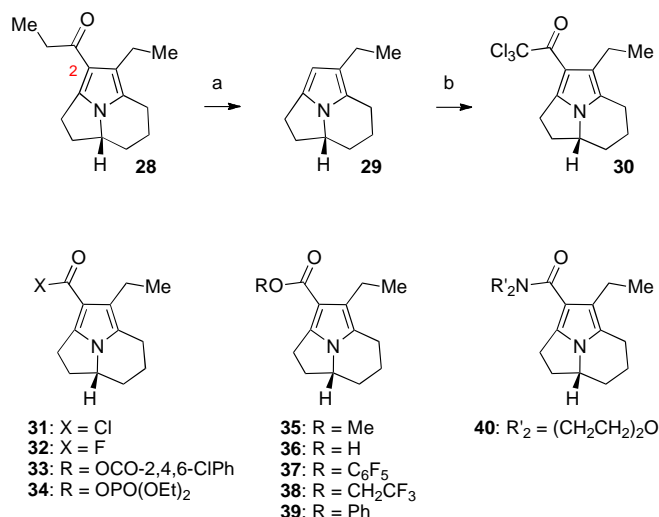
To address the sensitivity of these heterodimers we considered symmetrical diketone **27** as a versatile hexacyclic structure (Scheme 5). We anticipated that conjugation of an electron withdrawing carbonyl substituent to each of the pyrrole subunits would enhance the air stability of **27** and suppress undesired C3b alkylation. Importantly, preparation of β -substituted enone derivatives would be facilitated by enhanced acidity of the C2 methine in this structure, providing versatile substrates for radical, metal-mediated, and alternative electrophilic cyclization chemistries.⁹ Finally,



Scheme 5. Retrosynthesis of myrmicarin 430A (**4**) to heterodimer **27**.

the symmetry of this structure provided a powerful simplifying element, and enabled us to consider stereoselective introduction of the C3 ethyl substituent after heptacycle formation.

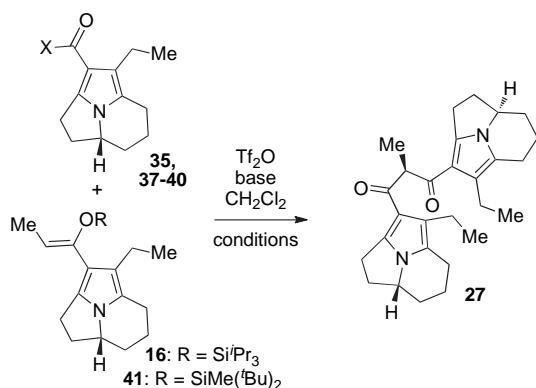
We envisioned that decarbonylation of readily available ketone **28**^{3b} and introduction of an appropriate acyl substituent via Friedel–Crafts reaction would provide the requisite truncated tricycle for preparation of **27** (Scheme 6). Accordingly, removal of the propanoyl group from **28** could be effected in high yield by heating a 1,2-dichloroethane solution of ketone **28** to 65 °C in the presence of trifluoromethanesulfonic acid and ethylene glycol. Treatment of pyrrole **29** with trichloroacetyl chloride in 1,2-dichloroethane at 65 °C in the absence of Lewis acid provided trichloroacetyl ketone **30** in 65% yield.¹⁰ Methanolysis of **30** proceeded efficiently to provide methyl ester **35**. Alternatively, hydrolysis of **30** and derivatization of acid **36** enabled preparation of a range of tricyclic acid, ester, and amides as electrophiles for heterodimerization with ketone **28**.



Scheme 6. Synthesis of acid, ester, and amide derivatives. Conditions: (a) TfOH, ethylene glycol, 1,2-dichloroethane, 65 °C, 1 h, 100%. (b) Cl_3CCOCl , 1,2-dichloroethane, 65 °C, 2 h, 74%.

Having established the efficiency of our trifluoromethanesulfonic anhydride mediated heterodimerization with ketone-substituted pyrroloindolizine electrophiles, we anticipated that this method would be applicable to the corresponding ester electrophiles. Unexpectedly, heterodimerization between methyl ester **35** and silyl enol ether **16** under previously optimized conditions provided diketone **27** in low yield. Raising the reaction temperature (Table 1, entries 1 and 2) decreased the yield of **27**. Use of an alternative base additive (Table 1, entry 3) resulted in decomposition of the silyl enol ether nucleophile. Modification of the silyl group of the nucleophile had little effect on the efficiency of the reaction (Table 1, entries 1 and 4). Use of electrophilic ester derivatives **37** ($\text{X}=\text{OC}_6\text{F}_5$), **38** ($\text{X}=\text{OCH}_2\text{CF}_3$), and **39** ($\text{X}=\text{OPh}$) and more Lewis basic morpholine amide **40** provided none of the desired product (Table 1, entries 5–8). Failure of these substrates to undergo heterodimerization may be a result of the reduced electrophilicity of the activated esters and amides relative to the corresponding ketone derivatives, which prevents addition of the silyl enol ether. The ability to access a range of activated acid and ester derivatives from pyrrolyl carboxylic acid **36** (Scheme 6) motivated us to investigate their direct reaction with a metalloenolate of ketone **28**. Surprisingly, the only products of attempted addition of the lithium enolate of **28** to the acid chloride **31**, acid fluoride **32**, or anhydride **33** were recovered **28** and tricyclic acid **36**. Efforts to add the potassium or lithium enolates of **28** to the activated ester derivatives **37**, **38**, or **39**

Table 1
Trifluoromethanesulfonic anhydride-promoted fragment coupling

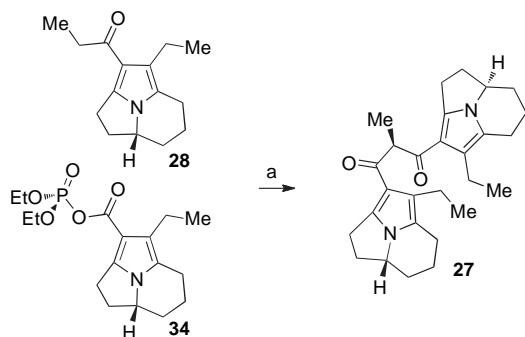


Entry	X (35 , 37–40)	Nucleophile	Base	Temp (°C)	Time (h)	Yield (%)
1	OMe	16	DTBMP	-78→0	7.5	30
2	OMe	16	DTBMP	-30→40	1.5	18
3	OMe	16	2-ClPyr	-78	0.5	0 ^a
4	OMe	41	DTBMP	-78→0	4	20
5	OC_6F_5	41	DTBMP	-78→0	3	0
6	OCH_2CF_3	16	DTBMP	-78→23	3	0
7	OPh	41	DTBMP	-78→0	5	0
8	$\text{N}(\text{CH}_2\text{CH}_2)_2\text{O}$	41	DTBMP	-78→0	2	0

^a Immediate decomposition of **16**. DTBMP=2,6-di-*tert*-butyl-4-methylpyridine; 2-ClPyr=2-chloropyridine.

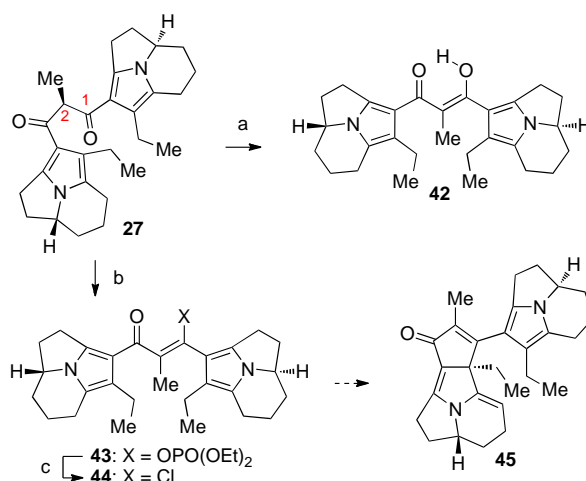
in the presence of various solvents and additives at different reaction temperatures failed to provide the desired heterodimer. The conversion of hydrolytically robust ester derivatives to the corresponding acid suggested that the metalloenolates of **28** may be undergoing O-acylation in preference to C-acylation and hydrolysis upon aqueous work up. In line with this hypothesis, we envisioned that the phosphate ester **34** (Scheme 6)¹¹ may provide strong coordination to the lithium cation of the corresponding enolate of **28** and promote a closed transition state to favor C-acylation. Gratifyingly, treatment of a toluene solution of ketone **28** with lithium bis(trimethylsilyl)amide at -78°C followed by addition of phosphate ester **34** and slow warming to -40°C furnished diketone **27** in 70% yield (Scheme 7). No product was observed under identical conditions when tetrahydrofuran, a coordinating solvent, was employed as solvent. As anticipated, hexacyclic diketone **27** demonstrated greatly enhanced stability relative to previous hexacyclic dimeric structures and could be stored for extended periods as a white solid. Interestingly, ^1H and ^{13}C spectra of diketone **27** in benzene- d_6 show two distinct sets of signals, which may reflect restricted rotation in the C1–C2 linkage and a preference for each pyrrole unit to maintain conjugation with its carbonyl substituent.

An efficient synthesis of diketone **27** enabled us to explore preparation of β -functionalized enone substrates for alternative



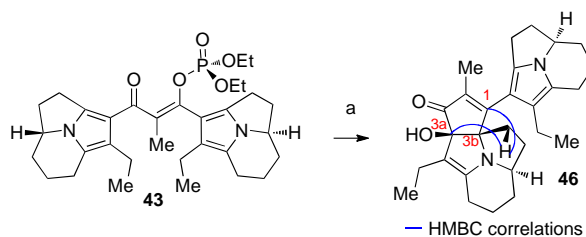
Scheme 7. Synthesis of heterodimer **27**. Conditions: (a) **28**, LHMDS, THF, -78°C , 15 min; **34**, $-78 \rightarrow -40^\circ\text{C}$, 1 h, 70%.

cyclization strategies. We anticipated that deprotonation of the C2 methine and trapping of the resulting enolate could provide vinyl trifluoromethanesulfonate, *p*-toluenesulfonate, or phosphate derivatives.⁶ Importantly, these conditions avoided electrophilic activation at C1 and provided useful substrates for a range of cyclization protocols. Quantitative formation of the symmetrical enol **42**¹² upon treatment of a tetrahydrofuran solution of diketone **27** with potassium bis(trimethylsilyl)amide followed by water quench verified the propensity of **27** to undergo enolization (Scheme 8). Attempts to trap the corresponding potassium or lithium enolate as a vinyl trifluoromethanesulfonate or *p*-toluenesulfonate, or its conversion to a vinyl selenide were unsuccessful. However, sequential treatment of diketone **27** in tetrahydrofuran with lithium bis(trimethylsilyl)amide and diethyl cyanophosphonate at -78°C efficiently provided the vinyl phosphate **43**. Interestingly, heating a dimethylformamide solution of phosphate **43** saturated with lithium chloride at 75°C for 4 h provided the vinyl chloride **44** as an alternative cyclization substrate.¹³



Scheme 8. Synthesis of enol **42** and β -substituted enone derivatives. Conditions: (a) KHMDS, THF, $-78 \rightarrow 0^\circ\text{C}$, 15 min; H_2O , 100%. (b) LHMDS, $(\text{EtO})_2\text{PCN}$, THF, -78°C , 1.2 h, 67%. (c) LiCl (satd), DMF, 75°C , 4 h, 43%.

The ability to prepare **44** via substitution of phosphate from **43** suggested that we may be able to improve the yield of **44** and apply this method to the synthesis of more reactive cyclization substrates. However, attempts to prepare **44** and the corresponding vinyl bromide derivative by activation of the C1 carbonyl and addition of halide typically resulted in hydrolysis to diketone **27** with poor mass recovery. Close monitoring of the reaction mixture upon treatment of **43** with triphenylphosphine dichloride in acetonitrile at 0°C revealed that diketone **27** was produced concomitantly with another highly fragile compound, which likewise underwent conversion to **27**. Careful purification of this acid, base, and air-sensitive compound and full structural characterization using two-dimensional NMR analysis revealed it to be the heptacyclic alcohol **46** (Scheme 9). Formation of **46** is consistent with enone activation



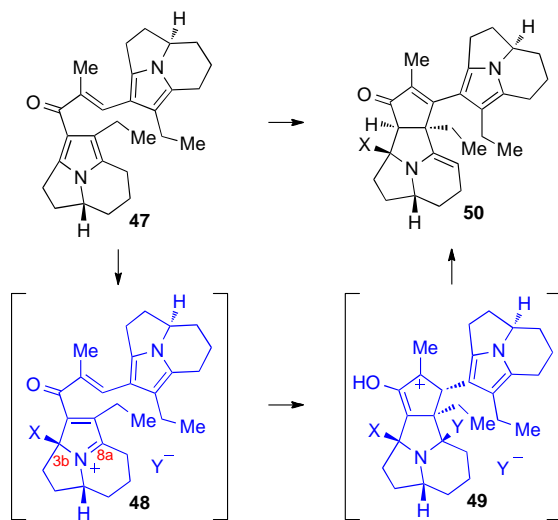
Scheme 9. Acid-promoted cyclization of **43**. Conditions: (a) Ph_3PCl_2 , CH_3CN , 0°C , 2 h, 36%.

followed by intramolecular pyrrole alkylation at C3b. Interestingly, hydrative trapping occurred exclusively at C3a, potentially as a consequence of strain relief upon sp^2 to sp^3 rehybridization at the junction between the two five-membered rings.

Successful preparation of the air stable diketone **27** and the corresponding β -substituted enone derivatives significantly expanded our knowledge regarding these dimeric pyrroloindolizine structures. These observations, combined with our expanded arsenal of methods to prepare heterodimeric structures, inspired strategies based on alternative electrophilic cyclization reactions.

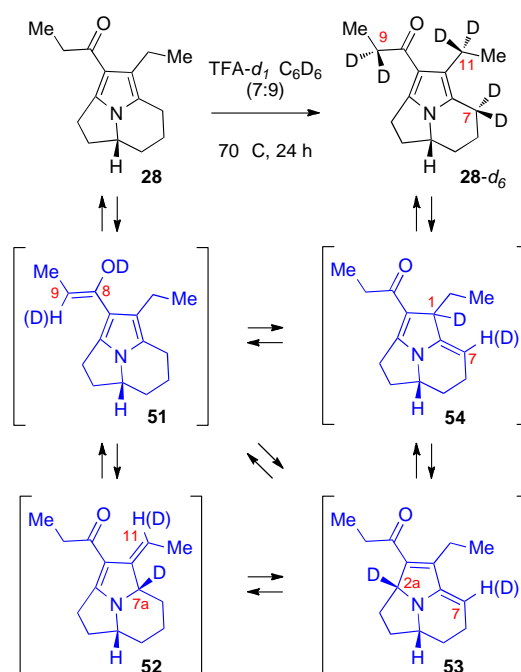
4. Design and synthesis of heterodimers for alternative electrophilic cyclization reactions

The propensity of the pyrroloindolizine subunit to undergo electrophilic trapping at C3b in dimeric structures suggested that we may be able to circumvent undesired alkylation by engaging this pyrrole nucleus in a reversible interaction with an alternative electrophile. Careful design of the dimeric structure may enable cyclization pathways in activated pyrrole substrates where C1–C3b bond formation could not occur. Specifically, we envisioned that reversible protonation, halogenation, or coordination of a metal salt at C3b would generate a dienone substructure such as **48**, which may be susceptible to a Nazarov cyclization to give heptacycle **50** (Scheme 10).¹⁴



Scheme 10. Potential activation and Nazarov cyclization of **47**.

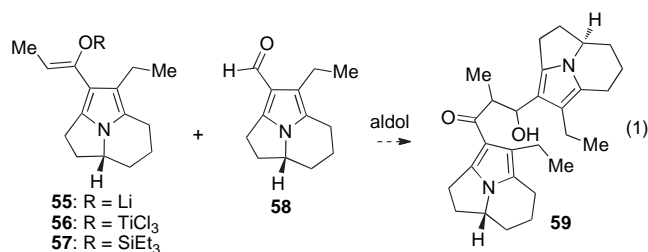
To examine the acid-promoted reactivity of a tricyclic substrate related to our proposed Nazarov cyclization we used in situ ^1H NMR monitoring to investigate protonation and deuterium incorporation in ketone **28** upon treatment with excess trifluoroacetic acid- d_1 (TFA- d_1).¹⁵ ^1H NMR analysis of the reaction mixture showed that addition of TFA- d_1 to a solution of **28** in benzene- d_6 (50% v/v) produced a 9:2 mixture of two protonated compounds. Heating this solution to 70 °C for 24 h resulted in deuterium incorporation at C9, but also at C7 and C11 (Scheme 11). Deuterium incorporation at C9 is consistent with tautomerization of the C8 carbonyl to the transient C8–C9 enol **51** and reversible protonation at C9 by TFA- d_1 . Additionally, H/D exchange at C7 and C11 suggests an equilibrium involving protonation of the pyrrole ring. Protonation at C7a followed by deprotonation at C11 may generate the vinyl enamine **52**, whereupon protonation by TFA- d_1 would result in deuterium incorporation at C11. Protonation at C2a and/or C1 would result in D incorporation at C7 via **53** and/or **54**. The inability to detect the proposed enamine intermediates by ^1H NMR indicates that they are



Scheme 11. Deuterium incorporation in **28** in the presence of excess TFA- d_1 .

present in exceedingly low concentrations relative to related ring protonated pyrrolium derivatives.

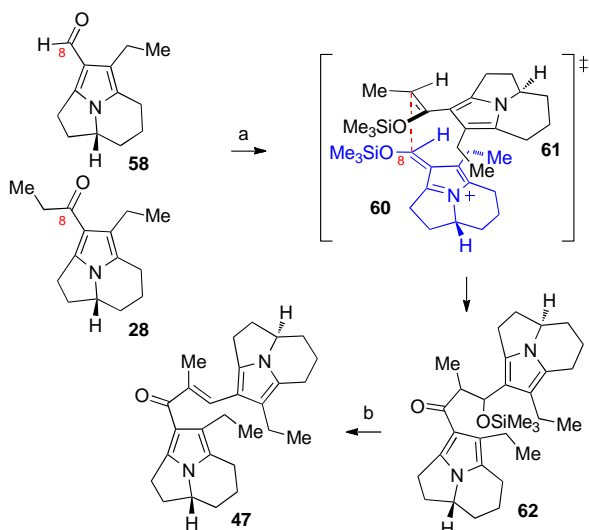
With evidence for reversible ring protonation of ketone **28** we aimed to exploit this reactivity to affect Nazarov cyclization in a dimeric structure. We envisioned that an aldol addition reaction between ketone **28** and tricyclic aldehyde **58**¹⁶ followed by elimination of the C1 alcohol could provide the enone substrate **47** (Scheme 10). Surprisingly, attempts to add lithium enolate **55** to aldehyde **58** failed to provide β -hydroxy ketone **59**, returning both starting materials (Eq. 1). Likewise, treatment of a mixture of **28** and **58** with titanium tetrachloride and *N,N*-diisopropylethylamine to effect in situ formation of the titanium enolate **56** and addition to **58** yielded no dimeric products. Use of Mukaiyama conditions for addition of triethylsilyl enol ether **57** to aldehyde **58** upon activation by titanium tetrachloride was also unsuccessful.



Equation 1. Attempted aldol addition reaction for the synthesis of ketone **59**.

The low electrophilicity of aldehyde **58** and the potential for the dimeric β -alkoxy ketone **59** to undergo retro-aldol reaction suggested that successful conditions may entail activation of **58** with concomitant generation of a silylated alcohol in the product. The unique susceptibility of carbonyl substituted pyrroloindolizine structures to activation by trialkylsilyl trifluoromethanesulfonates and accessibility of silyl enol ether derivatives of ketone **28** indicated that a one-pot Mukaiyama aldol conditions may be optimal for our system.¹⁷ In the event, addition of trimethylsilyl trifluoromethanesulfonate to a solution of ketone **28** and aldehyde **58**

in the presence of *N,N*-diisopropylethylamine efficiently generated the desired β -silyloxy ketone **62** in 92% yield (Scheme 12). Attempts to remove the trimethylsilyl group by treatment of **62** with tetra-*n*-butylammonium fluoride resulted in fragmentation to starting components **28** and **58**, consistent with our hypothesis that the β -alkoxy ketone is subject to a retro-aldol fragmentation. Fortunately, direct treatment of a dichloromethane solution of β -silyloxy ketone **62** with titanium tetrachloride at -78°C afforded the desired enone **47** in 84% yield, precluding the need for desilylation. This two-step sequence provided efficient access to our requisite Nazarov substrate.

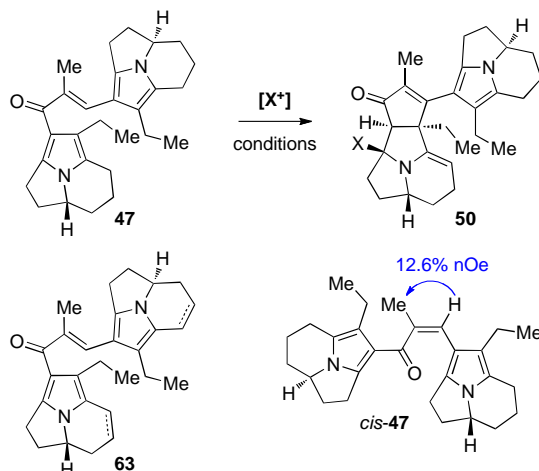


Scheme 12. Synthesis of enone **47**. Conditions: (a) TMSOTf, CH_2Cl_2 , 0°C , 30 min, 92%, 20: 13: 5: 5, mixture of diastereomers. (b) TiCl_4 , CH_2Cl_2 , -78°C , 25 min, 84%.

We considered a variety of electrophiles that may reversibly add to the pyrrole nucleus to generate a dienone substructure upon activation of **47**. In addition to use of acids to protonate **47** we envisioned that mild electrophilic halogenating agents or mercury (II) salts may reversibly add to C3b to generate **48** (Scheme 10). Interestingly, in situ monitoring showed that treatment of a benzene- d_6 solution of enone **47** with excess TFA- d_1 (50% v/v) and heating to 65°C afforded three major ring protonated species; however, no cyclization products could be detected, and basic quench returned starting enone **47** (Table 2, entry 1). By contrast, treatment of an acetonitrile–benzene (4:1) solution of **47** with the mild chlorinating agent 2,4,6,6-tetrachloro-2,4-cyclohexadienone (TCCHD) slowly generated a mixture of mono- and bis-alkene products **63** (Table 2, entry 2). This result is consistent with chlorination at C2a to form the desired intermediate for Nazarov cyclization (**48**, X=Cl, Scheme 10) followed by net elimination of hydrochloric acid to give the alkene(s) **63**. However, increased reaction temperatures resulted only in further oxidation and decomposition. Likewise, treatment of an acetonitrile- d_3 solution of enone **47** with mercury (II) trifluoroacetate resulted in rapid oxidation to **63** accompanied by isomerization to the corresponding enone *cis*-**47** (Table 2, entry 3). Exposure of **47** to mercury (II) acetate or mercury (II) chloride and heating to 65°C resulted only in partial isomerization (Table 2, entries 4 and 5).

We reasoned that nucleophilic additives might efficiently trap an initial cationic species **48** at C8a (Scheme 10) and increase the lifetime of the dienone substructure. Addition of saturating lithium chloride to an acetonitrile- d_3 –benzene- d_6 (4:1) solution of enone **47** in the presence of TFA (3.00 equiv) and heating to 65°C resulted in partial isomerization to the *Z*-isomer *cis*-**47**; however, no cyclization products were observed (Table 2, entry 6). Notably, no net

Table 2
Treatment of **47** with Brønsted and Lewis acids

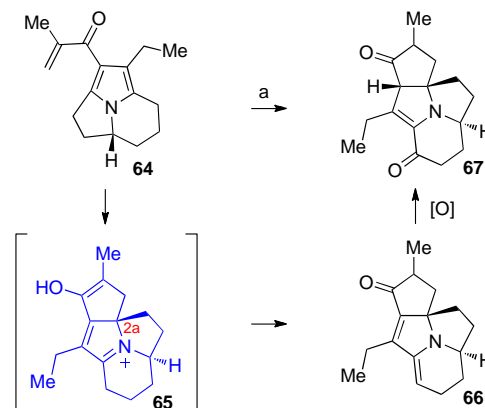


Entry	[X ⁺]	Additive	Solvent	Temp (°C)	Time (h)	Product (s)	(% yield)
1	TFA	—	C_6D_6	65	6	47	(67)
2	TCCHD	—	CD_3CN	23	4	63	(59)
3	$\text{Hg}(\text{O}_2\text{CCF}_3)_2$	—	CH_3CN	23	3	63	(57)
4	$\text{Hg}(\text{OAc})_2$	—	CD_3CN	65	7	47	(88), <i>cis</i> - 47 (4)
5	HgCl_2	—	CD_3CN	65	12	47	(47), <i>cis</i> - 47 (5)
6	TFA	LiCl	THF	65	4	47	(37), <i>cis</i> - 47 (8)
7	TFA	LiClO_4	THF	85	13	2	(20), 47 (30), <i>cis</i> - 47 (30)

TFA=trifluoroacetic acid; TCCHD=2,4,6,6-tetrachloro-2,4-cyclohexadienone.

isomerization occurred under these conditions in the absence of lithium chloride (Table 2, entry 1). Heating a solution of enone **47** and TFA (3.00 equiv) in tetrahydrofuran saturated with lithium perchlorate at 85°C gradually produced a 1:1 mixture of **47** and *cis*-**47** and resulted in fragmentation to myrmecarin 215B (**2**) (Table 2, entry 7). Use of alternative acids such as formic acid, trifluoromethanesulfonic acid, and perchloric acid resulted in varying degrees of fragmentation and isomerization, however, none of these conditions generated any observable cyclized products.

To investigate the feasibility of the Nazarov pathway in a simplified system containing only one pyrrole nucleus susceptible to activation we prepared the tricyclic enone substrate **64**. In contrast to our results with **47**, treatment of a dichloromethane solution of **64** with trifluoromethanesulfonic acid for 6 h followed by treatment 2-*tert*-butylimino-2-diethylamino-1,3-dimethyl-perhydro-1,3,2-diazaphosphorine (BEMP) yielded the tetracyclic product **67** (Scheme 13). Formation of this compound is consistent with protonation of the enone carbonyl followed by bond-formation at C2a.



Scheme 13. Acid-induced cyclization of **64**. Conditions: (a) TfOH, CH_2Cl_2 , 23°C , 6 h; BEMP, 23°C , 45 min, 16%.

Deprotonation at C7 would generate the enamine **66**, which may be susceptible to air oxidation upon exposure to atmosphere. In our earlier studies,^{3c} we observed that heptacyclic enamine intermediates similar to **66** were extremely sensitive to isolation due to rapid decomposition/oxidation pathways, rarely allowing isolation of identifiable oxidation products. In the present case, we were able to observe compound **66** by ¹H NMR and isolate the oxidation product **67** albeit in low yield.

Interestingly, while both tricyclic enone **64** and the β -functionalized enone **43** (Scheme 9) undergo C3b alkylation upon activation by Brønsted acid, enone **47** failed to show any evidence of analogous cyclization. The observation of ring-protonated derivatives of **47** by ¹H NMR and oxidation byproducts potentially arising via activation of the pyrrole ring provide evidence for formation of the desired Nazarov cyclization intermediate **48** (Scheme 10). Additionally, the failure of **47** to undergo C3b alkylation under these conditions provides validation for strategies that block undesired reactivity at this site similar to those described in the vinyl pyrrole series (Scheme 3). These studies on unique modes of pyrroloindolizine reactivity in dimeric structures inform the design of strategically functionalized heterodimers that are predisposed to differing reaction manifolds. Currently, we are pursuing the development of pyrroloindolizine derivatives that capitalize on these observations and are incorporating these derivatives into dimeric structures for the desired cyclization using the directed heterodimerization methods described herein.

5. Conclusion

Pyrroloindolizine structures undergo unique modes of electrophilic reactivity. Activation of C8-heteroatom substituted pyrroloindolizines and addition of neutral derivatives provides the corresponding dimeric structures. In this manner, we have employed trifluoromethanesulfonic anhydride activation of functionalized pyrroloindolizine substrates to prepare a range of hexacyclic structures. Studies on cyclopentannulation to form heptacyclic derivatives motivated the design of stable dimeric structures susceptible to further functionalization and derivatization. Accordingly, condensation between functional dimerization partners enabled us to prepare an air stable symmetrical diketone. Synthesis of a hexacyclic enone derivative through a uniquely efficient Mukaiyama aldol reaction allowed us to explore potential Nazarov cyclization in an activated dimeric structure. Our protocols for directed heterodimerization in these systems provide exciting opportunities to rapidly introduce the requisite functional groups in dimeric structures. In addition, fascinating observations on the reversible dimerization of myrmicarin 215B introduce the possibility of strategically modifying monomeric and dimeric pyrroloindolizine derivatives^{3d} to influence the structure of heptacyclic products formed under equilibrium conditions.¹⁸ The merger between strategic design of pyrroloindolizine structures and powerful methods of fragment assembly continues to guide our design of synthetic strategies relevant to these intriguing alkaloids and their derivatives.

6. Experimental section

6.1. General procedures

All reactions were performed in oven-dried or flame-dried round-bottomed flasks, modified Schlenk (Kjeldahl shape) flasks, or glass pressure vessels. The flasks were fitted with rubber septa and reactions were conducted under a positive pressure of argon. Stainless steel syringes or cannulae were used to transfer air- and moisture-sensitive liquids. Flash column chromatography was performed as described by Still and co-workers using silica gel

(60 Å pore size, 40–63 μ m, 4–6% H₂O content, Zeochem).¹⁹ Where necessary (so noted), silica gel was neutralized by treatment with the eluent containing 2.5% triethylamine. Analytical thin-layer chromatography was performed using glass plates pre-coated with 0.25 mm 230–400 mesh silica gel impregnated with a fluorescent indicator (254 nm). Where necessary (so noted), silica gel plates were neutralized by treatment with a solution of 2.5% triethylamine in ethyl acetate–hexanes followed by heating on a hot plate (~250 °C). Thin layer chromatography plates were visualized by exposure to ultraviolet light and/or by exposure to an aqueous solution of ceric ammonium molybdate (CAM) followed by heating (<1 min) on a hot plate (~250 °C). Organic solutions were concentrated on Büchi R-200 rotary evaporators at ~20 Torr at 25–35 °C unless otherwise indicated.

6.2. Materials

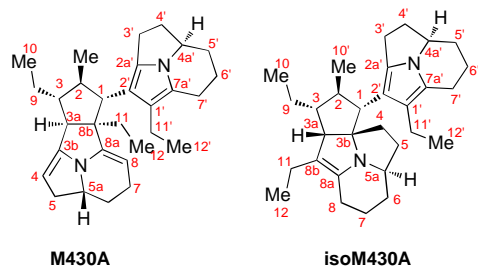
Commercial reagents and solvents were used as received with the following exceptions: dichloromethane, diethyl ether, tetrahydrofuran, acetonitrile, and toluene were purchased from J.T. Baker (Cyclotainer™) and were purified by the method of Grubbs and co-workers under positive argon pressure.²⁰ Triethylamine, diisopropylethylamine, benzene, and 1,2-dichloroethane were distilled over calcium hydride immediately before use. Ethylene glycol was distilled over magnesium sulfate at reduced pressure. Trichloroacetyl chloride was distilled at reduced pressure immediately before use. Trimethylsilyl trifluoromethanesulfonate was distilled at reduced pressure immediately before use.

6.3. Instrumentation

Proton nuclear magnetic resonance (¹H NMR) spectra were recorded with a Varian 300 Mercury or a Varian inverse probe 500 INOVA spectrometer or a Bruker 400 spectrometer or a Bruker inverse probe 600 Avance spectrometer. Chemical shifts are recorded in parts per million on the δ scale and are referenced from the residual protium in the NMR solvent (C₆D₅H: δ 7.16). Data is reported as follows: chemical shift [multiplicity (s=singlet, d=doublet, t=triplet, q=quartet, m=multiplet, app=apparent, br=broad), coupling constant(s) in Hertz, integration, assignment]. Carbon-13 nuclear magnetic resonance (¹³C NMR) spectra were recorded with a Varian 500 INOVA spectrometer and are recorded in parts per million on the δ scale and are referenced from the carbon resonances of the solvent (benzene-*d*₆: δ 128.4). Phosphorus-31 nuclear magnetic resonance (³¹P NMR) spectra were recorded with a Varian 500 INOVA spectrometer and are referenced to phosphoric acid (H₃PO₄: δ 0.00) as an external standard. Infrared data were obtained with a Perkin–Elmer 2000 FTIR and are reported as follows: [frequency of absorption (cm⁻¹), intensity of absorption (s=strong, m=medium, w=weak, br=broad), assignment]. We are grateful to Dr. Li Li for obtaining the mass spectrometric data at the Department of Chemistry's Instrumentation Facility, Massachusetts Institute of Technology. High-resolution mass spectra (HRMS) were recorded on a Bruker APEX 4.7 Tesler FTMS spectrometer using electrospray ion source (ESI) or electrospray (ES).

6.4. Positional numbering system

The numbering system for proton and carbon assignments for all pyrroloindolizine structures is consistent with the isolation reports for the naturally occurring tricyclic myrmicarins.^{1a} For direct comparison, the numbering system for proton and carbon assignments for all heterodimeric structures is consistent with the isolation reports for myrmicarin 430A^{1b} and the numbering system of isomyrmicarin 430A.^{3c}



6.5. Experimental procedures

6.5.1. Vinyl bromide 15. A solution of 2,4,4,6-tetrabromo-2,5-cyclohexadienone (TBCHD, 222 mg, 542 μmol , 2.00 equiv) in tetrahydrofuran (5.00 mL) was added to solution of tricyclic ketone **28**^{3b} (62.6 mg, 271 μmol , 1 equiv) in tetrahydrofuran (5.00 mL) at 0 °C over 10 min. After an additional 10 min the reaction mixture was concentrated under reduced pressure at 30 °C. The resulting green-brown oil was purified by flash column chromatography (silica gel pretreated with triethylamine: diam. 2.5 cm, ht. 15 cm, eluent: 1% Et₃N and 14% EtOAc in hexanes) to afford vinyl bromide **15** (56.4 mg, 68%) as a yellow solid. ¹H and ¹³C signals were assigned with the aid of gCOSY, HSQC, gHMBC, and NOE analysis. ¹H NMR (500 MHz, C₆D₆, 20 °C): 6.68 (d, *J*=2.9 Hz, 1H, C7–H), 3.34–3.27 (m, 1H, C4a–H), 2.96 (q, *J*=7.4 Hz, 2H, C11–H, C11–H'), 2.45–2.34 (m, 2H, C9–H, C9–H'), 2.32 (dd, *J*=16.0, 8.3 Hz, 1H C3–H_t), 2.26 (dd, *J*=16.4, 5.5 Hz, 1H, C5–H_c), 2.16 (ddd, *J*=16.3, 9.9, 6.8 Hz, 1H, C3–H_c), 2.02 (ddd, *J*=16.1, 12.6, 3.2 Hz, 1H, C5–H_t), 1.62 (dt, *J*=12.4, 6.2 Hz, 1H, C4–H_c), 1.39 (t, *J*=7.4 Hz, 3H, C12–H), 1.27 (t, *J*=7.3 Hz, 3H, C10–H), 1.18–1.10 (m, 1H, C4–H_t). ¹³C NMR (125 MHz, C₆D₆, 20 °C): 194.5 (C8), 139.8 (C2a), 128.8 (C1), 121.2 (C7a), 121.0 (C7), 117.8 (C2), 113.8 (C6), 53.3 (C4a), 41.4 (C5), 36.0 (C4), 34.7 (C9), 29.0 (C3), 20.0 (C11), 16.7 (C12), 9.1 (C10). FTIR (thin film) cm⁻¹: 2922 (m), 1650 (s), 1488 (m), 1401 (w), 1264 (w), 1097 (m). HRMS (ESI): calcd for C₁₅H₁₉BrNO [M+H]⁺: 308.0645, found: 308.0657. TLC (silica gel, 20% EtOAc in hexanes), *R*_f: 0.50 (UV, CAM).

6.5.2. Hexacyclic ketone 17. Trifluoromethanesulfonic anhydride (Tf₂O, 51.0 μL , 228 μmol , 2.00 equiv) was added portion-wise (3 \times 20 μL , 30 min intervals) to a solution of vinyl bromide **15** (47.0 mg, 152 μmol , 1 equiv), silyl enol ether **16**^{3a} (70.9 mg, 183 μmol , 1.20 equiv) (dried as a mixture by concentration from benzene under reduced pressure, 3 \times 1.00 mL), and 2,6-di-*tert*-butyl-4-methylpyridine (DTBMP, 157 mg, 570 μmol , 5.00 equiv) in dichloromethane (2.20 mL) at –78 °C. After 30 min, saturated aqueous sodium bicarbonate solution (2 mL) was added and the mixture was diluted with dichloromethane (5 mL) and allowed to warm to 23 °C. The aqueous layer was extracted with dichloromethane (4 \times 10 mL). The combined organic layers were dried over anhydrous sodium sulfate, were filtered, and were concentrated under reduced pressure. The resulting oil was purified by flash column chromatography (neutral alumina gel: diam. 3.0 cm, ht. 21 cm; eluent: 10% EtOAc in hexanes) to afford hexacyclic ketone **17** (61.6 mg, 78%, 9:2 mixture of diastereomers) as a yellow foam. ¹H and ¹³C signals for the major diastereomer were assigned with the aid of gCOSY, HSQC, gHMBC, and gNOESY analysis. ¹H NMR (500 MHz, C₆D₆, 20 °C): 6.82 (d, *J*=2.6 Hz, 1H, C8–H), 6.02–5.96 (m, 1H, C9–H), 3.99 (q, *J*=7.0 Hz, 1H, C2–H), 3.59–3.45 (m, 1H, C5a–H), 3.12 (dq, *J*=13.6, 7.0 Hz, 1H, C11'–H_x), 3.07–2.82 (m, 2H, C11'–H_y, C4a'–H), 2.72–2.52 (m, 1H, C3'–H_c), 2.58 (dd, *J*=15.4, 8.1 Hz, 1H, C4–H_x), 2.51–2.31 (m, 5H, C3'–H_t, C4–H_y, C11–H_x, C11–H_y, C7'–H_c), 2.29–2.20 (m, 1H, C6–H_y), 2.17–2.04 (m, 2H, C7'–H_t, C6–H_x), 1.90–1.66 (m, 5H, C10–H, C5–H_x, C4'–H_c), 1.60 (d, *J*=7.1 Hz, 3H, C10'–H), 1.54–1.47 (m, 1H, C6'–H_t), 1.43 (t, *J*=7.4 Hz, 3H, C12'–H), 1.40–1.27 (m, 3H, C5–H_y, C4'–H_t, C5'–H_c), 1.17–1.08

(m, 4H, C12–H, C6'–H_c), 0.66 (tdd, *J*=13.1, 10.8, 2.3 Hz, 1H, C5'–H_t). ¹³C NMR (125 MHz, C₆D₆, 20 °C): 196.3 (C1), 137.0 (C2a'), 136.8 (C3), 132.4 (C3b), 126.4 (C1'), 124.7 (C9), 122.1 (C8), 120.9 (C7a'), 119.6 (C8b), 117.9 (C2'), 115.6 (C3a), 111.4 (C8a), 111.2 (C7), 56.0 (C4a'), 53.3 (C5a), 51.9 (C2), 41.8 (C6), 37.4 (C5), 36.2 (C4'), 29.7 (C5'), 28.4 (C3'), 26.9 (C4), 22.7 (C6'), 20.2 (C11), 20.0 (C7'), 19.7 (C11'), 18.1 (C10'), 17.0 (C12), 16.3 (C12'), 15.7 (C10). HRMS (ESI): calcd for C₃₀H₃₈BrN₂O [M+H]⁺: 521.2162, found: 521.2164. TLC (silica gel, 20% EtOAc in hexanes), *R*_f: 0.39 (UV, CAM).

6.5.3. Diene 21. To a solution of hexacyclic alcohol **18** (17.3 mg, 33.0 μmol , 1 equiv) in benzene (1.30 mL) at 23 °C was added acetic acid (25.0 μL , 437 μmol , 13.0 equiv). After 1.8 h, saturated aqueous sodium bicarbonate solution (2 mL) was added and the mixture was diluted with dichloromethane (5 mL). The layers were separated and the aqueous layer was extracted with dichloromethane (4 \times 5 mL). The combined organic layers were dried over anhydrous sodium sulfate, were filtered, and were concentrated under reduced pressure. The resulting residue was purified by flash column chromatography (neutral alumina gel: diam. 2 cm, ht. 15 cm; eluent: 12.5% EtOAc in hexanes) to afford diene **21** (11.3 mg, 68%) as a yellow oil. Due to the air- and acid-sensitivity of **21**, best spectra were obtained by immediate spectroscopic analysis. ¹H and ¹³C signals were assigned with the aid of gCOSY, HSQC, and gHMBC analysis. ¹H NMR (500 MHz, C₆D₆, 20 °C): 6.88 (d, *J*=2.8 Hz, 1H, C8–H), 6.76 (s, 1H, C1–H), 6.04 (m, 1H, C9–H), 3.68–3.57 (m, 1H, C5a–H), 3.30 (tdd, *J*=10.6, 5.1, 3.8 Hz, 1H, C4a'–H), 2.64 (ddd, *J*=14.9, 10.9, 6.5 Hz, 1H, C3'–H_c), 2.62–2.39 (m, 7H, C7'–H_c, C11–H_x, C11–H_y, C11'–H_x, C11'–H_y, C4–H_x, C4–H_y), 2.46 (dd, *J*=15.0, 8.2 Hz, 1H, C3'–H_t), 2.40–2.29 (m, 2H, C7'–H_t, C6–H_x), 2.25 (s, 3H, C10'–H), 2.26–2.21 (m, 1H, C6–H_y), 1.94 (dt, *J*=11.4, 5.8 Hz, 1H, C4'–H_c), 1.86 (d, *J*=6.8 Hz, 3H, C10–H), 1.75–1.67 (m, 1H, C5–H_x), 1.67–1.60 (m, 1H, C6'–H_t), 1.50 (dq, *J*=12.4, 3.3 Hz, 1H, C5'–H_c), 1.48–1.40 (m, 2H, C5–H_y, C4'–H_t), 1.37–1.26 (m, 1H, C6'–H_c), 1.22–1.17 (m, 6H, C12'–H, C12–H), 0.77 (tdd, *J*=13.0, 10.9, 2.2 Hz, 1H, C5'–H_t). ¹³C NMR (125 MHz, C₆D₆, 20 °C): 140.7 (C2), 132.9 (C3a), 132.8 (C3b), 129.4 (C2a'), 127.9 (C1'), 123.3 (C1), 122.8 (C8b), 122.2 (C8) 121.5 (C9), 120.0 (C8a), 119.5 (C7a'), 114.5 (C3), 114.1 (C2'), 111.2 (C7), 55.6 (C4a'), 53.3 (C5a), 42.0 (C6), 37.8 (C4'), 37.8 (C5), 30.4 (C5'), 27.8 (C3'), 26.5 (C4), 22.8 (C6'), 20.8 (C7') 19.8 (C11'), 19.4 (C11), 17.1 (C12'), 17.0 (C12), 16.8 (C10'), 16.4 (C10). HRMS (ESI): calcd for C₃₀H₃₈BrN₂ [M+H]⁺: 505.2213, found: 505.2233. TLC (silica gel, 20% EtOAc in hexanes), *R*_f: 0.57 (UV, CAM).

6.5.4. Heptacyclic ketone 24. Trichloroacetic acid solution (0.16 M in benzene, 3.95 mg, 24.2 μmol , 1.05 equiv) was added portion-wise (3 \times 50 μL , 30 min intervals) to a solution of diene **21** (12.0 mg, 23.0 μmol , 1 equiv) in benzene (1.50 mL) at 23 °C. After 1.8 h, saturated aqueous sodium bicarbonate solution (2 mL) was added and the mixture was diluted with dichloromethane (3 mL). The layers were separated and the aqueous layer was extracted with dichloromethane (4 \times 4 mL). The combined organic layers were dried over anhydrous sodium sulfate, were filtered, and were concentrated under reduced pressure. The resulting brown oil was purified by flash column chromatography (silica gel pretreated with triethylamine: diam. 1.0 cm, ht. 15 cm; eluent: 1% Et₃N and 14–99% EtOAc in hexanes) to afford heptacyclic ketone **24** (5.00 mg, 48%) as a yellow oil. Due to the fragility and air-sensitivity of **24**, best spectra were obtained by immediate spectroscopic analysis. ¹H and ¹³C signals were assigned with the aid of gCOSY, gHSQC, gHMBC, and gNOESY analysis. ¹H NMR (500 MHz, C₆D₆, 20 °C): 5.60 (s, 1H, C8–H), 3.52–3.48 (m, 1H, C5a–H), 3.34 (s, 1H, C1–H), 3.08–3.02 (m, 1H, C4a'–H), 2.65 (dd, *J*=17.7, 7.5 Hz, 1H, C6–H_y), 2.63–2.54 (m, 4H, C7'–H_c, C11'–H, C11'–H', C3'–H_c), 2.41–2.33 (m, 1H, C7'–H_x), 2.37 (d, *J*=17.7 Hz, 1H, C6–H_x), 2.29–2.12 (m, 5H, C3'–H_t, C11–H, C11–H', C9–H, C9–H'), 2.05

(ddd, $J=12.9, 10.7, 2.1$ Hz, 1H, C4–H_x), 1.98–1.88 (m, 2H, C4'–H_c, C5–H_x), 1.70 (dt, $J=11.6, 8.8$ Hz, 1H, C4–H_y), 1.60 (s, 3H, C10'–H), 1.58–1.48 (m, 2H, C4'–H_t, C6'–H_t), 1.42–1.36 (m, 2H, C5'–H_c, C5–H_y), 1.31 (t, $J=7.5$ Hz, 3H, C12'–H), 1.21–1.18 (m, 1H, C6'–H_c), 1.08 (t, $J=7.6$ Hz, 3H, C12–H), 0.99 (t, $J=7.6$ Hz, 3H, C10–H), 0.80–0.72 (m, 1H, C5'–H_t). ¹³C NMR (125 MHz, C₆D₆, 20 °C): 189.6 (C7), 177.0 (C8a), 164.2 (C3a), 152.4 (C2), 134.1 (C3), 127.5 (C2a'), 124.5 (C8b), 121.8 (C1'), 118.3 (C7a'), 111.4 (C2'), 97.9 (C8), 88.0 (C3b), 58.5 (C5a), 55.8 (C1), 55.0 (C4a'), 41.1 (C4), 38.4 (C6), 37.0 (C4'), 31.2 (C5), 29.9 (C5'), 26.7 (C3'), 22.8 (C6'), 21.1 (C7'), 19.6 (C11'), 19.1 (C11), 18.4 (C9), 17.6 (C12), 15.1 (C10'), 14.6 (C10), 14.4 (C12). FTIR (thin film) cm⁻¹: 2856 (s), 1704 (m), 1643 (m), 1453 (m), 1206 (m), 1061 (m). HRMS (ESI): calcd for C₃₀H₃₉N₂O [M+H]⁺: 443.3057, found: 443.3076. TLC (1.0% Et₃N in EtOAc, Et₃N neutralized silica gel), R_f: 0.55 (UV, CAM).

6.5.5. Pyrrole 29. To a vigorously stirred solution of ketone **28**^{3b} (124.0 mg, 536 μmol, 1 equiv) and ethylene glycol (600 μL, 10.7 mmol, 20.0 equiv) in 1,2-dichloroethane (5.25 mL) at 23 °C was added trifluoromethanesulfonic acid (142 μL, 1.61 mmol, 3.00 equiv) and the resulting yellow solution was heated to 65 °C. After 1 h, the tan colored reaction mixture was allowed to cool to 23 °C, was poured into 3 N KOH (17.5 mL), and was diluted with diethyl ether (20 mL). The layers were separated and the aqueous layer was extracted with diethyl ether (3×20 mL). The combined organic layers were washed with brine (20 mL), were dried over anhydrous sodium sulfate, were filtered, and were concentrated under reduced pressure. The resulting material was purified by flash column chromatography (silica gel pretreated with triethylamine: diam. 2.5 cm, ht. 15 cm; eluent: 2% Et₃N and 3% EtOAc in hexanes) to afford pyrrole **28** (93.9 mg, 100%) as white needles. The spectroscopic data was consistent with the data reported in the literature.²¹ ¹H NMR (500 MHz, C₆D₆, 20 °C): 5.95 (s, 1H, C2–H), 3.26 (tdd, $J=10.7, 4.8, 3.8$ Hz, 1H, C4a–H), 2.65–2.55 (m, 4H, C3–H_c, C3–H_t, C11–H, C11'–H), 2.52 (dd, $J=14.5, 8.1$ Hz, C7–H_c), 2.38 (ddd, $J=16.8, 11.3, 6.1$ Hz, 1H, C7–H_t), 1.94 (dt, $J=11.5, 5.8$ Hz, 1H, C4–H_c), 1.69–1.64 (m, 1H, 6–H_t), 1.58–1.48 (m, 2H, C4–H_t, C5–H_c), 1.40–1.30 (m, 1H, C6–H_c), 1.32 (t, $J=7.5$ Hz, 3H, C12–H), 0.82 (tdd, $J=13.0, 10.6, 2.5$ Hz, 1H, C5–H_t). ¹³C NMR (125 MHz, C₆D₆, 20 °C): 130.7, 123.5, 118.3, 100.3, 55.5, 37.8, 30.3, 25.9, 23.1, 20.9, 16.8. FTIR (thin film) cm⁻¹: 2958 (s), 1425 (m), 1322 (m), 1168 (w), 767 (m). HRMS (ESI): calcd for C₁₂H₁₈N [M+H]⁺: 176.1434, found: 176.1443. TLC (silica gel pretreated with Et₃N, 2.5% Et₃N in [17.5% EtOAc in hexanes]), R_f: 0.69 (UV, CAM).

6.5.6. Trichloroacetylpyrrole 30. To a solution of pyrrole **29** (68.8 mg, 393 μmol, 1 equiv) in 1,2-dichloroethane (4.00 mL) at 23 °C was added trichloroacetyl chloride (65.9 μL, 590 μmol, 1.50 equiv) and the solution was heated to 65 °C. After 2 h, the resulting deep red reaction mixture was allowed to cool to 23 °C, was poured into saturated aqueous sodium bicarbonate solution (10 mL), and was diluted with ethyl acetate (10 mL). The layers were separated and the aqueous layer was extracted with ethyl acetate (3×12.5 mL). The combined organic layers were washed with brine (10 mL), were dried over anhydrous sodium sulfate, were filtered, and were concentrated under reduced pressure. The resulting material was purified by flash column chromatography (silica gel pretreated with triethylamine: diam. 2.0 cm, ht. 23 cm; eluent: 2% Et₃N and 3% EtOAc in hexanes) to afford trichloroacetylpyrrole **30** (93.5 mg, 74%) as a pale yellow solid. ¹H NMR (500 MHz, C₆D₆, 20 °C): 3.03 (dd, $J=16.2, 7.8$ Hz, 1H, C3–H_t), 2.95–2.80 (m, 3H, C11–H, C11'–H, C4a–H), 2.49 (ddd, $J=16.5, 10.8, 6.1$ Hz, 1H, C3–H_c), 2.26 (dd, $J=16.6, 6.3$ Hz, 1H, C7–H_c), 1.99 (ddd, $J=16.7, 12.0, 6.7$ Hz, 1H, C7–H_t), 1.68 (dt, $J=11.7, 5.9, 5.9$ Hz, 1H, C4–H_c), 1.48–1.41 (m, 1H, C6–H_t), 1.35 (t, $J=7.4$ Hz, 3H, C12–H), 1.35–1.25 (m, 2H, C4–H_t, C5–H_c), 1.10–1.00 (m, 1H, C6–H_c), 0.57

(tdd, $J=13.2, 11.0, 2.3$ Hz, 1H, C5–H_t). ¹³C NMR (125 MHz, C₆D₆, 20 °C): 177.2, 138.3, 129.8, 122.0, 107.2, 99.1, 56.5, 36.0, 31.6, 29.2, 22.2, 20.5, 19.7, 15.8. FTIR (thin film) cm⁻¹: 2958 (s), 1666 (s), 1456 (m), 1319 (m), 813 (m). HRMS (ESI): calcd for C₁₄H₁₇Cl₃NO [M+H]⁺: 320.0370, found: 320.0358. TLC (silica gel pretreated with Et₃N, 2.5% Et₃N in [2.5% EtOAc in hexanes]), R_f: 0.22 (UV, CAM).

6.5.7. Methyl ester 35. To a solution of trichloroacetylpyrrole **30** (40.0 mg, 125 μmol, 1 equiv) in methanol–dichloromethane (15:1, 3.20 mL) at 23 °C was added sodium methoxide (135.1 mg, 2.50 mmol, 20.0 equiv) in a single portion. After 1.5 h, aqueous ammonium chloride solution (10 mL) was added and the reaction mixture was diluted with ethyl acetate (12.5 mL). The layers were separated and the aqueous layer was extracted with ethyl acetate (4×7.5 mL). The combined organic layers were washed with brine (10 mL), were dried over anhydrous sodium sulfate, were filtered, and were concentrated under reduced pressure. The resulting material was purified by flash column chromatography (silica gel: diam. 1.5 cm, ht. 18 cm; eluent: 7.5→15% EtOAc in hexanes) to afford methyl ester **35** (25.2 mg, 87%) as a white solid. ¹H NMR (500 MHz, C₆D₆, 20 °C): 3.65 (s, 3H, OCH₃), 3.11–2.92 (m, 4H, C11–H, C11'–H, C3–H_t, C4a–H), 2.70 (ddd, $J=16.2, 10.4, 6.3$ Hz, 1H, C3–H_c), 2.40 (ddd, $J=16.3, 6.5, 1.1$ Hz, 1H, C7–H_c), 2.18 (ddd, $J=16.4, 11.9, 6.8$ Hz, 1H, C7–H_t), 1.78 (ddd, $J=11.2, 6.2, 5.6$ Hz, 1H, C4–H_c), 1.56–1.50 (m, 1H, C6–H_t), 1.46 (t, $J=7.4$ Hz, 3H, C12–H), 1.41–1.32 (m, 2H, C4–H_t, C5–H_c), 1.21–1.11 (m, 1H, C6–H_c), 0.66 (tdd, $J=13.1, 11.0, 2.3$ Hz, 1H, C5–H_t). ¹³C NMR (125 MHz, C₆D₆, 20 °C): 166.2, 138.8, 125.6, 120.3, 106.8, 56.1, 50.4, 36.4, 29.8, 27.5, 22.7, 20.2, 19.8, 16.6. FTIR (thin film) cm⁻¹: 2965 (s), 1702 (s), 1514 (m), 1435 (m), 1275 (m), 1110 (m). HRMS ESI calcd for C₁₄H₂₀NO₂ [M+H]⁺: 234.1489, found: 234.1496. TLC (silica gel, 10% EtOAc in hexanes), R_f: 0.25 (UV, CAM).

6.5.8. Acid phosphate 34. To a vigorously stirred suspension of acid **36** (40.3 mg, 183 μmol, 1 equiv) and diethyl cyanophosphonate (39.7 μL, 265 μmol, 1.45 equiv) in ethyl acetate (1.83 mL) at 23 °C was added triethylamine (34.4 μL, 247 μmol, 1.35 equiv). After 1.5 h, the reaction mixture was concentrated under reduced pressure to a volume of approximately 100 μL and the resulting material was immediately purified by flash column chromatography (silica gel pretreated with triethylamine: diam. 1.5 cm, ht. 15 cm; eluent: 2.5% Et₃N and 67.5% EtOAc in hexanes) to afford acid phosphate **34** (55.7 mg, 85%) as a pale yellow oil. Acid phosphate **34** was sensitive to hydrolysis and was used immediately after preparation. ¹H NMR (500 MHz, C₆D₆, 20 °C): 4.30–4.20 (m, 4H, OCH₂CH₃), 2.95–2.78 (m, 4H, C11–H, C11'–H, C4a–H, C3–H_t), 2.57 (ddd, $J=16.3, 10.4, 6.2$ Hz, 1H, C3–H_c), 2.27 (dd, $J=16.6, 3.8$ Hz, 1H, C7–H_c), 2.04 (ddd, $J=16.5, 11.8, 6.7$ Hz, 1H, C7–H_t), 1.69 (dt, $J=11.7, 6.0$ Hz, 1H, C4–H_c), 1.50–1.44 (m, 1H, C6–H_t), 1.36 (t, $J=7.5$ Hz, 3H, C12–H), 1.30 (dq, $J=12.5, 3.4$ Hz, 1H, C5–H_c), 1.27–1.19 (m, 1H, C4–H_t), 1.10–1.02 (m, 7H, OCH₂CH₃, C6–H_c), 0.58 (tdd, $J=13.1, 11.0, 2.3$ Hz, 1H, C5–H_t). ¹³C NMR (125 MHz, C₆D₆, 20 °C): 158.9 (d, $J=7.7$ Hz), 141.5, 126.7 (d, $J=1.4$ Hz), 121.3, 105.1 (d, $J=9$ Hz), 64.6 (d, $J=5.4$ Hz), 64.6 (d, $J=5.4$ Hz), 56.4, 35.9, 29.5, 27.8, 22.5, 20.0, 19.5, 16.6 (d, $J=6.8$ Hz), 16.6 (d, $J=6.8$ Hz), 16.2. ³¹P NMR (200 MHz, C₆D₆, 20 °C): -6.14. FTIR (thin film) cm⁻¹: 2965 (m), 1726 (s), 1513 (m), 1273 (s), 1031 (s), 1004 (s). HRMS (ESI): calcd for C₁₇H₂₆NNaO₅P [M+Na]⁺: 378.1441, found: 378.1443. TLC (silica gel pretreated with Et₃N, 2.5% Et₃N in [47.5% EtOAc in hexanes]), R_f: 0.37 (UV, CAM).

6.5.9. Diketone 27. A solution of ketone **28**^{3b} (63.2 mg, 273 μmol, 1.76 equiv) in toluene (300 μL) was transferred via cannula to a solution of lithium hexamethyldisilazide (LHMDS, 97% w/w, 46.2 mg, 268 μmol, 1.60 equiv) in toluene (550 μL) at -78 °C and the transfer was completed with a toluene rinse (2×100 μL). The reaction flask was warmed to 0 °C for 15 min then cooled to -78 °C for 5 min. A solution of acid phosphate **34** (55.7 mg, 155 μmol, 1 equiv) in toluene

(300 μL) was added via cannula and transfer was complete with a toluene rinse ($2 \times 100 \mu\text{L}$). The reaction flask was placed on a dry-ice–acetone bath at -40°C and allowed to warm slowly to 0°C . After 1 h, aqueous ammonium chloride solution (7.5 mL) was added and the reaction mixture was diluted with ethyl acetate (12.5 mL). The layers were separated and the aqueous layer was extracted with ethyl acetate ($4 \times 10 \text{ mL}$). The combined organic layers were washed with brine (10 mL), were dried over anhydrous sodium sulfate, were filtered, and were concentrated under reduced pressure. The resulting material was purified by flash column chromatography (silica gel: diam. 2.0 cm, ht. 17 cm; eluent: $40 \rightarrow 55\%$ EtOAc in hexanes, followed by 1% AcOH in 99% EtOAc for elution of acid **36**) to afford diketone **27** (47.0 mg, 70%) as a white solid. In addition, ketone **28** (28.3 mg, 38% recovered) and acid **36** (8.1 mg, 16% recovered based on acid phosphate **34**) were separated during chromatography and isolated. ^1H and ^{13}C signals were assigned with the aid of gCOSY, HSQC, and gHMBC analysis. ^1H NMR (500 MHz, C_6D_6 , 20°C): 4.45 (m, 1H, C2–H), 3.18–2.98 (m, 3H, C11–H_x, C11–H_y, C4a'–H), 3.05 (q, $J=7.4$ Hz, 2H, C11'–H_x, C11'–H_y), 2.97–2.89 (m, 1H, C5a–H), 2.94 (dd, $J=15.4, 8.2$ Hz, 1H, C4–H_t), 2.78–2.66 (m, 2H, C3'–H_c, C3'–H_t), 2.58 (ddd, $J=15.5, 10.6, 6.5$ Hz, C4–H_c), 2.40–2.29 (m, 2H, C8–H_c, C7'–H_c), 2.17–2.07 (m, 2H, C8–H_t, C7'–H_t), 1.81–1.75 (m, 2H, C5–H_c, C4'–H_c), 1.78 (d, $J=7.2$ Hz, 3H, C10'–H), 1.53–1.36 (m, 4H, C7–H_c, C6'–H_c, C5–H_t, C4'–H_t), 1.51 (t, $J=7.4$ Hz, 3H, C12–H), 1.45 (t, $J=7.4$ Hz, 3H, C12'–H), 1.36–1.29 (m, 2H, C6–H_c, C5'–H_c), 1.20–1.10 (m, 1H, C6'–H_t), 1.12–1.02 (m, 1H, C7–H_t), 0.68 (tdd, $J=14.8, 13.2, 2.3$ Hz, 1H, C6–H_t), 0.59 (tdd, $J=14.8, 13.3, 2.3$ Hz, 1H, C5'–H_t). ^{13}C NMR (125 MHz, C_6D_6 , 20°C): 193.3 (C1), 193.0 (C3), 136.2 (C3b), 136.2 (C2a'), 126.7 (C8b), 126.5 (C1'), 121.1 (C7a'), 120.9 (C8a), 117.6 (C3a), 117.6 (C2'), 56.6 (C2), 56.0 (C4a'), 55.9 (C5a), 36.5 (C5), 36.3 (C4'), 29.7 (C6), 29.4 (C5'), 28.6 (C3'), 28.1 (C4), 22.7 (C6'), 22.5 (C7), 20.4 (C8), 20.3 (C7'), 20.0 (C11), 19.9 (C11'), 16.4 (C12), 16.3 (C12'), 15.8 (C10'). FTIR (thin film) cm^{-1} : 2928 (s), 1650 (s), 1494 (m), 1428 (m), 1320 (w), 972 (s). HRMS (ESI): calcd for $\text{C}_{28}\text{H}_{36}\text{N}_2\text{NaO}_2$ [$\text{M}+\text{Na}$] $^+$: 455.2669, found: 455.2664. TLC (silica gel, 50% EtOAc in hexanes), R_f : 0.35 (UV, CAM).

6.5.10. Vinyl phosphate 43. A solution of diketone **27** (17.1 mg, 39.5 μmol , 1 equiv) in tetrahydrofuran (200 μL) was added drop wise to a solution of potassium hexamethyldisilazide (KHMDs, 16.6 mg, 79.1 μmol , 2.00 equiv) in tetrahydrofuran (550 μL) at -78°C . The reaction flask was warmed to 0°C for 15 min then cooled to -78°C for 5 min and diethyl cyanophosphonate (13.1 μL , 79.1 μmol , 2.00 equiv) was added. After 1.2 h, water was added (6 mL) and the reaction mixture was diluted with ethyl acetate (10 mL) and brine (1 mL) and allowed to warm to 23°C . The layers were separated and the aqueous layer was extracted with ethyl acetate ($4 \times 7.5 \text{ mL}$). The combined organic layers were washed with brine (7 mL), were dried over anhydrous sodium sulfate, were filtered, and were concentrated under reduced pressure. The resulting material was purified by flash column chromatography (silica gel: diam. 1.5 cm, ht. 20 cm; eluent: $50 \rightarrow 85\%$ EtOAc in hexanes) to afford vinyl phosphate **43** (15.0 mg, 67%) as a pale yellow oil. ^1H and ^{13}C signals were assigned with the aid of gCOSY and gHMBC analysis. ^1H NMR (500 MHz, C_6D_6 , 20°C): 3.90–3.65 (m, 4H, OCH_2CH_3), 3.57–3.30 (m, 3H, C4–H_x, C4–H_y, C3'–H_x), 3.30–3.20 (m, 2H, C4a'–H, C5a–H), 3.22–3.12 (m, 1H, C11'–H_x), 3.00 (dq, $J=13.7, 7.1$ Hz, 1H, C11'–H_y), 2.83–2.71 (m, 2H, C3'–H_y, C11–H_x), 2.66 (dq, $J=14.6, 7.4$ Hz, 1H, C11–H_y), 2.50 (dd, $J=16.2, 6.1$ Hz, 1H, C8–H_c), 2.43 (dd, $J=16.5, 5.6$ Hz, 1H, C7'–H_c), 2.27 (ddd, $J=16.8, 11.3, 6.1$ Hz, 1H, C8–H_t), 2.22–2.09 (m, 2H, C7'–H_t, C4'–H_c), 2.06 (d, $J=2.6$ Hz, 3H, C10'–H), 2.06–2.00 (m, 1H, C5–H_c), 1.83–1.65 (m, 2H, C4'–H_t, C5–H_t), 1.65–1.48 (m, 3H, C7–H_t, C6'–H_t, C6–H_c), 1.52 (t, $J=7.4$ Hz, 3H, C12'–H), 1.45–1.40 (m, 1H, C5'–H_c), 1.37 (t, $J=7.6$ Hz, 3H, C12–H), 1.30 (m, 1H, C7–H_c), 1.20 (m, 1H, C6'–H_c), 0.93 (td, $J=7.1, 0.8$ Hz, 3H, OCH_2CH_3), 0.86 (td, $J=7.1, 0.9$ Hz, 3H, OCH_2CH_3), 0.70–0.62 (m, 2H, C5'–H_t, C6–H_t). ^{13}C NMR (125 MHz, C_6D_6 , 20°C):

191.9 (d, $J=0.9$ Hz, C3), 140.8 (C2a'), 139.1 (d, $J=6.9$ Hz, C1), 132.2 (C3b), 127.2 (d, $J=9.6$ Hz, C2), 125.7 (C1'), 123.5 (C8b), 121.1 (C7a'), 119.1 (C8a), 116.7 (C2'), 108.6 (C3a), 63.7 (d, $J=5.5$ Hz, OCH_2CH_3), 63.4 (d, $J=5.5$ Hz, OCH_2CH_3), 56.4 (C4a'), 55.9 (C5a), 37.5 (C5), 36.6 (C4'), 30.4 (C6), 29.9 (C5'), 27.9 (C3'), 26.3 (C4), 23.0 (C6'), 22.8 (C7), 21.0 (C11'), 20.2 (C8), 20.2 (C7'), 19.8 (C11), 17.1 (C10'), 16.5 (d, $J=2.3$ Hz, OCH_2CH_3), 16.4 (d, $J=2.3$ Hz, OCH_2CH_3), 16.4 (C12'), 15.9 (C12). ^{31}P NMR (200 MHz, C_6D_6 , 20°C): -4.33 . FTIR (thin film) cm^{-1} : 2931 (s), 1704 (s), 1621 (s), 1503 (m), 1267 (m), 1091 (m), 1047 (m). HRMS (ESI): calcd for $\text{C}_{32}\text{H}_{46}\text{N}_2\text{O}_5\text{P}$ [$\text{M}+\text{H}$] $^+$: 569.3139, found: 569.3137. TLC (silica gel, 65% EtOAc in hexanes), R_f : 0.15 (UV, CAM).

6.5.11. Heptacyclic alcohol 46. Triphenylphosphine dichloride (PPh_3Cl_2 , 6.5 mg, 20.0 μmol , 1.50 equiv) was added as solid in one portion to a solution of vinyl phosphate **43** (7.4 mg, 13.0 μmol , 1 equiv) in acetonitrile (280 μL) at 0°C . After 2 h, saturated aqueous sodium bicarbonate solution (3 mL) was added and the mixture was diluted with ethyl acetate (5 mL). The layers were separated and the aqueous layer was extracted with ethyl acetate ($2 \times 5 \text{ mL}$). The combined organic layers were dried over anhydrous sodium sulfate, were filtered, and were concentrated under reduced pressure. The resulting residue was purified by flash column chromatography (silica gel pretreated with triethylamine: diam. 1.0 cm, ht. 15 cm; eluent: 2.5% Et_3N and 7.5% EtOAc in hexanes) to afford heptacyclic alcohol **46** (2.0 mg, 36%) as a pale yellow oil. Due to the fragility and air-, acid-, and base-sensitivity of **46**, best spectra were obtained by immediate spectroscopic analysis. ^1H and ^{13}C signals were assigned with the aid of gCOSY, gHSQC, gHMBC, and gNOESY analysis. ^1H NMR (500 MHz, C_6D_6 , 20°C): 3.51–3.45 (m, 1H, C4a'–H), 3.16–3.11 (m, 1H, C5a–H), 3.02 (br s, 1H, OH), 2.89 (ddd, $J=15.5, 10.9, 6.2$ Hz, 1H, C3'–H_c), 2.73 (dd, $J=15.2, 7.9$ Hz, 1H, C3'–H_t), 2.57 (dd, $J=16.5, 6.8$ Hz, 1H, C7'–H_c), 2.45–2.32 (m, 7H, C4–H_x, C11–H, C11–H', C11'–H, C11'–H', C7'–H_t, C8–H_x), 2.06–1.99 (m, 2H, C4'–H_c, C8–H_y), 1.95 (s, 3H, C10'–H), 1.93–1.86 (m, 1H, C4–H_y), 1.72–1.62 (m, 2H, C6'–H_t, C5–H_x), 1.60–1.28 (m, 7H, C5'–H_c, C6–H_x, C6–H_y, C4'–H_t, C6'–H_c, C7–H_c, C7–H_t), 1.25–1.19 (m, 1H, C5–H_y), 1.21 (t, $J=7.5$ Hz, 3H, C12–H), 1.16 (t, $J=7.5$ Hz, 3H, C12'–H), 0.82 (m, 1H, C5'–H_t). ^{13}C NMR (125 MHz, C_6D_6 , 20°C): 206.8 (C3), 165.3 (C1), 145.4 (C8a), 136.0 (C2), 129.8 (C2a'), 122.9 (C1'), 120.2 (C7a'), 115.1 (C8b), 109.3 (C2'), 89.3 (C3a), 83.2 (C3b), 62.5 (C5a), 56.1 (C4a'), 37.8 (C4'), 33.0 (C6), 30.1 (C5'), 29.5 (C4), 29.0 (C5), 27.5 (C3'), 24.5 (C8), 22.7 (C6'), 21.2 (C7), 21.1 (C7'), 19.8 (C11'), 17.9 (C11), 16.6 (C12), 16.3 (C12'), 10.9 (C10'). FTIR (thin film) cm^{-1} : 3445 (br m), 2936 (s), 1687 (s), 1628 (s), 1452 (m), 1315 (m), 1046 (m). HRMS (ESI): calcd for $\text{C}_{28}\text{H}_{37}\text{N}_2\text{O}_2$ [$\text{M}+\text{H}$] $^+$: 433.2850, found: 433.2857. TLC (2.5% Et_3N in [12.5% EtOAc in hexanes], Et_3N neutralized silica gel), R_f : 0.20 (UV, CAM).

6.5.12. Hexacyclic enone 47. To a solution of β -trimethylsilyloxy ketone **62** (32.6 mg, 64.3 μmol , 1 equiv) in dichloromethane (1.25 mL) at -78°C was added titanium tetrachloride (1.0 M in dichloromethane, 96.5 μL , 96.5 μmol , 1.50 equiv). After 30 min, brine was added (4 mL) and the reaction mixture was diluted with ethyl acetate (6 mL) and allowed to warm to 23°C . The layers were separated and the aqueous layer was extracted with ethyl acetate ($3 \times 5 \text{ mL}$). The combined organic layers were washed sequentially with saturated aqueous sodium bicarbonate solution (4 mL) and brine (4 mL), were dried over anhydrous sodium sulfate, were filtered, and were concentrated under reduced pressure. The resulting material was purified by flash column chromatography (silica gel pretreated with triethylamine: diam. 2.0 cm, ht. 14 cm; eluent: 2.5% Et_3N and 20–30% EtOAc in hexanes) to afford enone **47** (22.4 mg, 84%) and *cis*-enone *cis*-**47** (1.4 mg, 5%) as yellow solids. ^1H and ^{13}C signals for enone **47** were assigned with the aid of gCOSY, HSQC, and gHMBC analysis. ^1H signals for *cis*-enone *cis*-**47** were assigned with the aid of gCOSY analysis. Data for **47**: ^1H NMR

(500 MHz, C_6D_6 , 20 °C): 7.64 (q, $J=1.4$ Hz, 1H, C1–H), 3.30 (dq, $J=13.9$, 7.1 Hz, 1H, C11– H_x), 3.26–3.20 (m, 1H, C5a–H), 3.16–3.08 (m, 1H, C4a'–H), 3.00 (dq, $J=13.6$, 7.4 Hz, C11– H_y), 2.75–2.52 (m, 3H, C3'– H_c , C3'– H_t , C4– H_c), 2.50 (d, $J=1.3$ Hz, 3H, C10'–H), 2.50–2.35 (m, 5H, C4– H_t , C8– H_c , C11'– H_x , C11'– H_y , C7'– H_c), 2.32–2.19 (m, 2H, C8– H_t , C7'– H_t), 1.92 (dt, $J=11.7$, 5.9 Hz, 1H, C5– H_c), 1.72 (dt, $J=11.7$, 5.9 Hz, 1H, C4'– H_c), 1.63–1.51 (m, 2H, C7– H_t , C6'– H_t), 1.53 (t, $J=7.4$ Hz, 3H, C12–H), 1.52–1.32 (m, 4H, C5– H_t , C6– H_c , C5'– H_c , C4'– H_t), 1.32–1.12 (m, 2H, C7– H_c , C6'– H_c), 1.09 (t, $J=7.6$ Hz, 3H, C12'–H), 0.79–0.65 (m, 2H, C5'– H_t , C6– H_t). ^{13}C NMR (125 MHz, C_6D_6 , 20 °C): 194.1 (C3), 137.6 (C2a'), 134.3 (C2), 132.6 (C1), 131.0 (C3b), 126.0 (C8b), 123.5 (C1'), 121.1 (C8a), 120.2 (C7a'), 117.7 (C3a), 113.6 (C2'), 56.0 (C4a'), 55.8 (C5a), 37.8 (C5), 36.9 (C4'), 30.1 (C6), 30.0 (C5'), 28.9 (C3'), 28.5 (C4), 22.7 (C6'), 22.6 (C7), 20.7 (C7'), 20.3 (C8), 20.1 (C11), 19.1 (C11'), 16.9 (C12'), 16.6 (C12), 16.0 (C10'). FTIR (thin film) cm^{-1} : 2928 (s), 1607 (s), 1426 (m), 1320 (m), 1035 (w), 734 (w). HRMS (ESI): calcd for $C_{28}H_{37}N_2O$ [$M+H$] $^+$: 417.2900, found: 417.2913. TLC (silica gel pretreated with Et_3N , 2.5% Et_3N in [25% EtOAc in hexanes]), R_f : 0.25 (UV, CAM). Data for *cis*-**47**: 1H NMR (500 MHz, C_6D_6 , 20 °C): 6.54 (q, $J=1.4$ Hz, 1H, C1–H), 3.33–3.16 (m, 2H, C11– H_x , C11– H_y), 3.11 (dd, $J=15.8$, 8.2 Hz, 1H, C3'– H_t), 3.00–2.89 (m, 2H, C5a–H, C3'– H_c), 2.81 (dd, $J=15.9$, 8.1 Hz, 1H, C4– H_t), 2.77–2.61 (m, 4H, C4– H_c , C4a'–H, C11'– H_x , C11'– H_y), 2.47 (dd, $J=16.2$, 6.4 Hz, 1H, C8– H_c), 2.39 (dd, $J=16.6$, 6.1 Hz, 1H, C7'– H_c), 2.36–2.27 (m, 1H, C7'– H_t), 2.16 (ddd, $J=17.0$, 11.5, 6.3 Hz, 1H, C8– H_t), 1.86 (dt, $J=11.4$, 5.8 Hz, 1H, C5– H_c), 1.74 (td, $J=11.7$, 5.9 Hz, 1H, C4'– H_c), 2.25 (d, $J=1.3$ Hz, 1H, C10'–H), 1.62 (t, $J=7.4$ Hz, 3H, C12–H), 1.50–1.39 (m, 2H, C6'– H_t , C7– H_t), 1.38–1.31 (m, 2H, C4'– H_t , C5– H_t), 1.35 (t, $J=7.5$ Hz, 3H, C12'–H), 1.29–1.13 (m, 2H, C6– H_c , C5'– H_c), 1.05–0.89 (m, 2H, C6'– H_c , C7– H_c), 0.74–0.61 (m, 2H, C5'– H_t , C6– H_t). TLC (silica gel pretreated with Et_3N , 2.5% Et_3N in [25% EtOAc in hexanes]), R_f : 0.44 (UV, CAM).

6.5.13. Tetracycle 67. To a solution of tricyclic enone **64** (8.4 mg, 34.0 μ mol, 1 equiv) in dichloromethane (800 μ L) at 23 °C was added trifluoromethanesulfonic acid (6.1 μ L, 69.0 mmol, 2.00 equiv). After 6 h, 2-*tert*-butylimino-2-diethylamino-1,3-dimethyl-perhydro-1,3,2-diazaphosphorine on polystyrene (BEMP, 165 mg, 340 mmol, 10.00 equiv) was added. After 45 min, the reaction mixture filtered through a cotton plug and the orange filtrate was concentrated under reduced pressure. The resulting residue was purified by flash column chromatography (silica gel pretreated with triethylamine: diam. 1.0 cm, ht. 12 cm; eluent: 2.5% Et_3N and 12.5% EtOAc in hexanes) to afford tetracycle **67** (1.4 mg, 16%) as a yellow oil. 1H and ^{13}C signals were assigned with the aid of gCOSY, HSQC, and gHMBC analysis. 1H NMR (500 MHz, C_6D_6 , 20 °C): 3.29 (dq, $J=14.3$, 7.5, 0.8 Hz, 1H, C11– H_x), 3.02 (dq, $J=9.9$, 6.4 Hz, 1H, C4a–H), 2.94 (s, 1H, C2–H), 2.72 (dq, $J=14.5$, 7.4, 1.7 Hz, 1H, C11– H_y), 2.10–2.00 (m, 2H, C6– H_x , C9–H), 1.98 (td, $J=14.9$, 4.2 Hz, 1H, C6– H_y), 1.64 (dd, $J=13.4$, 5.3 Hz, 1H, C10– H_x), 1.55–1.47 (m, 2H, C4– H_x , C10– H_y), 1.42 (ddt, $J=13.6$, 5.9, 4.0 Hz, 1H, C5– H_x), 1.39–1.31 (m, 1H, C3– H_x), 1.19 (t, $J=7.5$ Hz, 3H, C12–H), 1.19–1.12 (m, 1H, C3– H_y), 1.03 (d, $J=7.4$ Hz, 3H, C10'–H), 0.99–0.80 (m, 2H, C4– H_y , C5– H_y). ^{13}C NMR (125 MHz, C_6D_6 , 20 °C): 216.4 (C8), 195.4 (C7), 138.2 (C7a), 126.2 (C1), 78.5 (C2a), 63.4 (C2), 58.1 (C4a), 44.4 (C9), 42.7 (C10'), 37.8 (C6), 36.4 (C3), 32.5 (C4), 27.1 (C5), 20.2 (C11), 17.1 (C10), 13.4 (C12). FTIR (thin film) cm^{-1} : 2962 (s), 1737 (s), 1692 (s), 1599 (m), 1453 (m), 1198 (m), 1032 (w). HRMS (ESI): calcd for $C_{16}H_{22}NO_2$ [$M+H$] $^+$: 260.1645, found: 260.1646. TLC (silica gel pretreated with Et_3N , 2.5% Et_3N in [12.5% EtOAc in hexanes]), R_f : 0.18 (UV, CAM).

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 - Subjecting of **43** to conditions described for Heck reaction of vinyl phosphates failed to effect cyclization, instead returning **43** or diketone **27**. Attempts to achieve intermolecular Heck reaction also failed to provide the corresponding products, consistent with the expectation that **43** would not be an optimal substrate for these transformations. Likewise, subjecting of **44** to conditions known to effect cross coupling with aryl chlorides did not induce Heck cyclization.
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